

Chapter 13 States Of Matter

A theory that explains the states of matter, based on the concept that all matter consists of tiny particles that are in constant motion Gas Pressure Results from the force exerted by a gas per unit surface area of an object; due to collisions of gas particles with the object

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A theory explaining the states of matter based on the concepts that all matter consists of tiny particles that are in constant motion.

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Chapter 13: States of Matter - Jayne Heier

Chapter 13 States of Matter pages 341 to 362. Properties of fluids. Gases and liquids are both fluids. Both these states of matter have greater freedom of motion.

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You are already familiar with the three common states of matter: solid, liquid, and gas. Solid objects litter the room around you. For example, you can easily recognize the

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Chapter 13 - States of Matter - 13.4 Changes of State - 13.4 Lesson Check - Page 439: 26 Answer they represent the pressure and temperature in which two phases exist in equilibrium.

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Chapter 13 States of Matter137 SECTION 13.1 THE NATURE OF GASES (pages 385–389) This section introduces the kinetic theory and describes how it applies to gases. It defines gas pressure and explains how temperature is related to the kinetic energy of the particles of a substance. Kinetic Theory and a Model for Gases (pages 385–386) 1.

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Chapter 13- The States of Matter Gases- indefinite volume and shape, low density. Liquids- definite volume, indefinite shape, and high density. Solids- definite volume and shape, high density

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